

<b>Question</b>	<b>Answer</b>	<b>Marks</b>	<b>AO Element</b>	<b>Notes</b>	<b>Guidance</b>
1(a)	decomposition	<b>1</b>			
1(b)	31.7 (g)	<b>1</b>			
2(a)	melting (1) condensing (1) freezing (1) sublimation (1)	<b>4</b>			
2(b)	no new substances are made <b>or</b> the change can be reversed (by a physical process)	<b>1</b>			
2(c)	boiling happens at a specific temperature <b>or</b> evaporation happens over a range of temperatures	<b>1</b>			
3	2 (CO <sub>2</sub> ) (1) 3 (H <sub>2</sub> O) (1)	<b>2</b>			

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4	<p><b>A</b> / substance <b>Y</b> dissolves easily in water (1)</p> <p><b>C</b> / substance <b>Y</b> melts (at 801 °C) (1)</p> <p>the change can be reversed by altering the conditions (1)</p>	<b>3</b>			
5	<p><b>C</b> / boils (at 1330°C) (1)</p> <p><b>D</b> / dissolves (readily in water) (1)</p> <p>the change can be reversed by altering the conditions (1)</p>	<b>3</b>			
6(a)	<p><b>M1</b> calcium oxide</p> <p><b>M2</b> CaO</p>	<b>2</b>			
6(b)	(step) <b>3</b>	<b>1</b>			
6(c)	thermal decomposition	<b>1</b>			
6(d)	heating	<b>1</b>			
6(e)	$\text{Ca(OH)}_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$	<b>1</b>			
6(f)	<p><b>M1</b> CO<sub>2</sub> is acidic</p> <p><b>M2</b> Ca(OH)<sub>2</sub> is a base / alkali</p>	<b>2</b>			

- Mark Scheme

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